T3 Examination Specifications (2016-2017)

Subject	Chemistry
Grade	11
Duration	90 minutes
Mark Determination	• One step or one idea is equivalent to at least one degree and half a score is not distributed to any question. • Each score must be consistent with the level of skill required in the question. • The student must receive one degree for each expected answer. • A score is set in parentheses, so [] for each parent or sub-grade at the end of the question. • Minimum one idea is only one degree.
Question Types/ Details	• Multiple-choice. • Short answers Questions. • Matching Questions. • Multiple particle questions, long answers or a combination of both.
Outcomes which may be selected for assessment in the exam	• The test will include 16 learning outcomes

CL2.M1.13	**Identify** the types of heterogeneous and homogeneous
CL2.M1.14	**Compare** the properties of suspensions, colloids, and
CL2.M1.15	**Describe** the electrostatic forces in the colloids described.
CL2.M1.16	**Describe** the concentration using different units.
CL2.M1.17	**Determine** the concentrations of solutions.
CL2.M1.18	**Calculate** the molarity of a solution.
CL2.M1.19	**Describe** the intermolecular forces affect solvation.
CL2.M1.20	**Define** solubility and describe the factors which affect
CL2.M1.21	**Define** energy and describe how potential and kinetic energy differ.
CL2.M1.22	**Explain** how chemical potential energy is related to the heat loss or gained in chemical reactions.
CL2.M1.23	**Calculate** amount of heat absorbed or released by a substance as its temperature changes.
CL2.M1.24	**Describe** how a calorimeter is used to measure energy that is absorbed or released.
CL2.M1.25	**Explain** the meaning of enthalpy and enthalpy change in chemical reactions and processes.
CL2.M1.26	**Write** the thermochemical equations for combustion, vaporization, and fusion.
CL2.M1.27	**Describe** how energy is lost or gained during changes of
CL2.M1.28	**Calculate** the heat that is absorbed or released in a chemical reaction calculated.